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The purpose of this paper is to examine the nature of chemical change in the hope of recognizing and setting forth the basic principles that help us to understand why they occur.

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Only one equilibrium constant for gas-phase reactions (Chapter 12), the thermodynamic constant  $K$ , often referred to as  $K_p$ . This simplifies not only the treatment of gaseous equilibrium but also the discussion of reaction spontaneity (Chapter 16) and electrochemistry (Chapter 17). 2.

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P S reactions are indicated.  $\Delta^\ddagger G^\ominus$  is the overall standard free-energy change in the direction S P.

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